

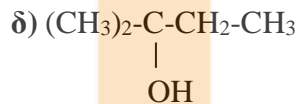
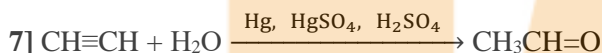
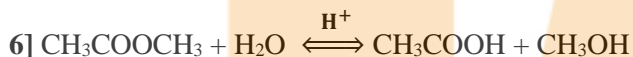
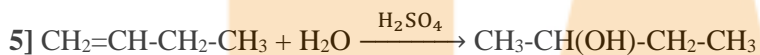
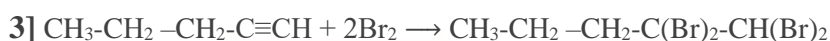
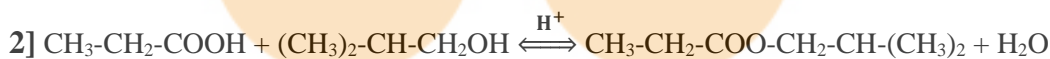
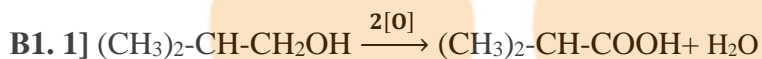
**ΕΝΔΕΙΚΤΙΚΕΣ ΑΠΑΝΤΗΣΕΙΣ ΔΙΑΓΩΝΙΣΜΑΤΟΣ
ΧΗΜΕΙΑΣ Β' ΛΥΚΕΙΟΥ**

Επιμέλεια διαγωνίσματος: ΑΓΓΕΛΑΚΟΠΟΥΛΟΣ ΜΑΡΙΝΟΣ -
ΓΡΗΓΟΡΟΠΟΥΛΟΣ ΔΗΜΗΤΡΗΣ - ΤΣΑΚΑΝΙΑ ΜΑΡΙΑ - ΦΡΑΣΕΡΙ ΜΑΡΙΝΑ

ΘΕΜΑ Α

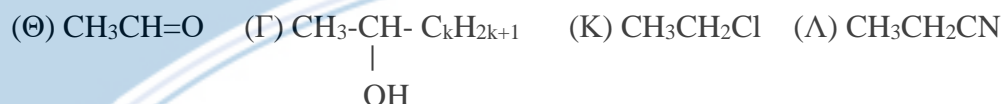
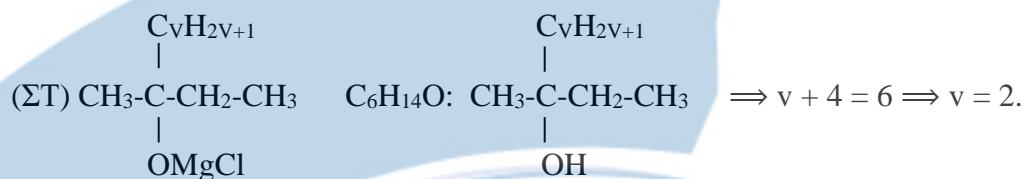
A1. β A2. γ A3. α A4. γ A5. β

ΘΕΜΑ Β



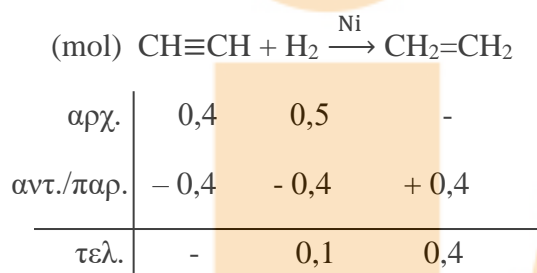
ΘΕΜΑ Γ



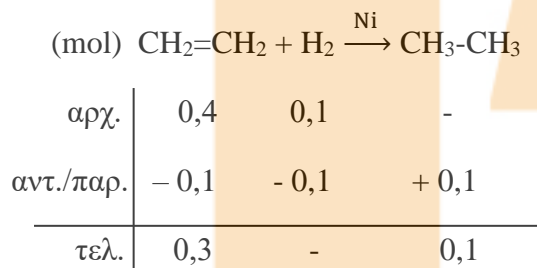


ΘΕΜΑ Δ

Δ1. α) $\text{CH}\equiv\text{CH: } n = \frac{m}{M_r} = \frac{10,4 \text{ gr}}{26 \frac{\text{gr}}{\text{mol}}} = 0,4 \text{ mol}, \quad \text{H}_2: n = \frac{m}{M_r} = \frac{1 \text{ gr}}{2 \frac{\text{gr}}{\text{mol}}} = 0,5 \text{ mol}$



Συνεχίζεται η υδρογόνωση:



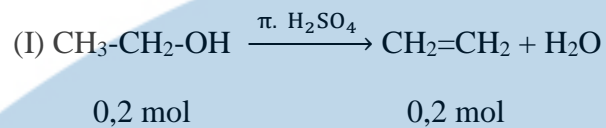
Τελική σύσταση: 0,3 mol $\text{CH}_2=\text{CH}_2$ και 0,1 mol $\text{CH}_3\text{-CH}_3$



0,3 mol 0,3 mol

$\text{Br}_2: C = \frac{n}{V} \Rightarrow V = \frac{n}{C} = \frac{0,3 \text{ mol}}{0,1 \frac{\text{mol}}{\text{L}}} = 0,3\text{L} = 300\text{mL}$

$$\Delta 2. \text{CH}_3\text{-CH}_2\text{-OH: } n = \frac{m}{M_r} = \frac{9,2 \text{ gr}}{46 \frac{\text{gr}}{\text{mol}}} = 0,2 \text{ mol}$$

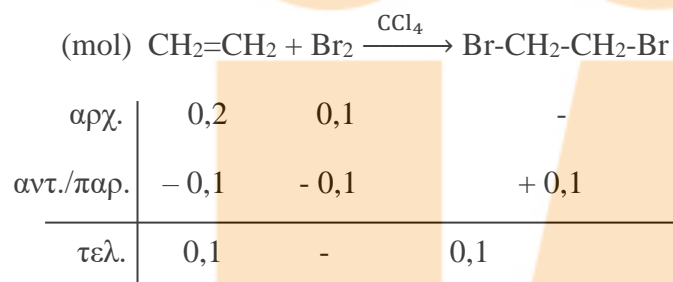


$$\text{(A) CH}_2=\text{CH}_2: m = n \cdot M_r = 0,2 \text{ mol} \cdot 28 \frac{\text{gr}}{\text{mol}} = 5,6 \text{ gr}$$

(II) Br₂: Σε 100mL διαλύματος Br₂/CCl₄ περιέχονται 4gr Br₂

Σε 400mL διαλύματος Br₂/CCl₄ περιέχονται x = 16gr Br₂

$$\Rightarrow n = \frac{m}{M_r} = \frac{16 \text{ gr}}{160 \frac{\text{gr}}{\text{mol}}} = 0,1 \text{ mol}$$



Δεν περισσεύει Br₂ ⇒ αποχρωματίζεται το διάλυμα Br₂/CCl₄.